

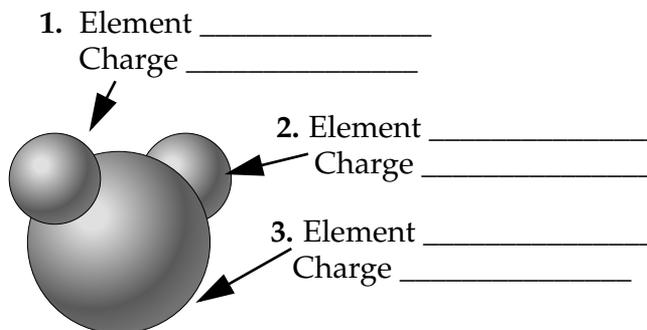
**Earth: The Water Planet** ▪ *Review and Reinforce*

# The Properties of Water

## Understanding Main Ideas

Label the parts of this water molecule by writing the name of the element and the electrical charge in items 1 through 3.

Answer the following questions on a separate sheet of paper.



4. Why is water considered a polar substance?
5. Which state of water allows fish to remain in a lake when winter temperatures are below 0°C? Explain.
6. What happens to the molecules of water vapor when the temperature of the gas cools to 100°C?
7. Why is water often called the “universal solvent”?

## Building Vocabulary

Match each term with its definition by writing the letter of the correct definition in the right column on the line beside the term in the left column.

- |                           |  |
|---------------------------|--|
| _____ 8. capillary action | a. a mixture that forms when one substance dissolves another   |
| _____ 9. condensation     | b. form of a substance, including solid, liquid, or gas  |
| _____ 10. evaporation     | c. the tightness across the surface of water caused by the polar molecules pulling on each other                 |
| _____ 11. specific heat   | d. the process by which molecules at the surface of a liquid absorb enough energy to change to the gaseous state |
| _____ 12. solution        | e. the combined force of attraction among water molecules and with the molecules of surrounding materials        |
| _____ 13. solvent         | f. the process by which a gas changes to a liquid  |
| _____ 14. state           | g. a substance that dissolves another substance  |
| _____ 15. surface tension | h. the amount of heat needed to increase the temperature of a certain amount of a substance by 1°C               |